

- ① metals - oxidize
 non-metal - reduce (24)
 - metals have a low number
 of valence electrons than non-metals
 (24)

- ② If the X^{n+} changes then
 oxidation or reduction has
 occurred

③

	a.	$O_2 \rightarrow O^0 + O^0$	state 0, 0
tough	b.	$H^+ SO_4^{2-}$	1, 6, -2
ones	c.	$Na^+ H^+ PO_4^{3-}$	1, 1, 5, -2
→	d.	$Cr^{+3} Cl_3^{-1}$	+3, -1
		(loses e^-)	(gains e^-)

④

a.	Fe oxidized	0 - Reduced
b.	Zn "	Ag - "
c.	K "	Cl - "
d.	Ca "	O - "